Nanocatalysts from Ionic Liquid Precursors for the Direct Conversion of CO₂ to Hydrocarbons

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Introduction

The direct conversion of carbon dioxide (CO₂) into lower olefins (C₂-C₄) is a highly desirable process as a sustainable production route. Lower olefins, i.e., ethylene, propylene and butenes (C₂-C₄), are key building blocks in the current chemical industry. The reaction proceeds *via* two main consecutive reactions: Reverse Water Gas Shift (RWGS) to produce CO followed by the further conversion of CO to hydrocarbons *via* the Fischer–Tropsch reaction². This process is achieved by a multifunctional iron-based catalyst supported on zeolites providing three types of active sites (Fe₃O₄, Fe₃C₂ and acid sites), which cooperatively catalyse a tandem reaction¹.

To date, attempts at synthesising a suitable catalyst for the direct hydrogenation reaction follow a conventional precipitation procedure, whereby Iron Oxide Nanoparticles (*IONs*) are produced and then embedded within a zeolite structure by granule mixing. This method provides limited control over the size and shape of the IONs formed; a characteristic of imperative importance due to its significant effect on the hydrocarbon product distribution obtained. In our novel approach, ionic liquids are utilised for the synthesis of the *IONs* resulting in better control over size and morphology of the nanostructured material, and as a consequence, better conversion and selectivity towards the olefins.

Materials and Methods

Na–Fe₃O₄ nanocatalysts obtained by a one-pot synthesis method, already demonstrated in literature², employed a precipitation method (**PM**) involving a mixture of iron (II) and iron (III) chloride hydrate, deionised water and HCl to result in a clear solution. Subsequently, NaOH has been added as a precipitating agent, which resulted in the formation of a black precipitate, which consisted of magnetite, Fe₃O₄ (**PM-Na**). To prepare an Fe₃O₄ sample without sodium, ammonium hydroxide was used as the precipitating agent. These samples were then used as a benchmark for the following novel method. An ionic liquid-assisted synthesis method consisted of heating two iron precursors of Fe (II) and Fe (III) chloride hydrates in a reaction medium of [C₄mim][OAc] ionic liquid (**IL**). After the synthesis, the Fe₃O₄ particle were embedded in either HZSM-5 (SiO₂/Al₂O₃ = 300) or HZSM-5 (SiO₂/Al₂O₃ = 80) zeolite by either granular mixing or mixing in a ball mill at a mass ratio of the two components of 1:1.

 CO_2 hydrogenation reactions were performed at 320 °C, 25 bar and a H/CO₂ ratio of 3:1, in a stainless steel fixed-bed reactor with an inner diameter of 15 mm. Typically, 0.75 g of catalyst Fe₃O₄/Zeolite was mixed with alumina (20–40 meshes) in a 1/1 mass ratio. Prior to reaction, the catalyst was reduced *in-situ* at 350 °C for 8 h in a pure H₂ flow at atmospheric pressure. All products from the reactor were analysed with an online gas chromatograph (GC) equipped with ShinCarbon and PONA columns.

The materials were characterized with ICP-OES (Agilent 5100 ICP-OES), PXRD (PANanalytical X'Pert Pro Diffractometer), TEM (G2 Talos) operated at 200 kV, H_2 Chemisorption (Micromeritics AutoChem II 2920).

Results and Discussion

The ionic-liquid assisted synthesis of a nanocrystalline magnetite precursor showed that ionic liquids provide a controlled precipitation method thanks to their dual functionality as solvent and templating agent. Characterization of the prepared catalysts with the IL method by PXRD shows high phase purity for magnetite, Fe_3O_4 , small particle size and TEM shows good dispersion with the zeolite component. The compounds obtained by ionic liquid method result in amorphous XRD patterns before calcination, however, show phase purity and good crystallinity after calcination at 420° C under a flow of N₂. This has also been confirmed by SEM and TEM (Figure 1.a, b). Hydrogen temperature-programmed reduction (H₂-TPR) was used to determine the reducibility of the Fe₃O₄ and the hydrogen uptake.

Hydrogen temperature-programmed reduction (H₂-TPR) was used to determine the interaction between Fe species and the support. As shown in Figure 1.c, all the catalysts present two peaks with increasing reduction temperature, which are assigned to the conversions Fe₃O₄—FeO and FeO—Fe, respectively. It is observed that catalyst (**PM**) starts to be reduced at lower temperature compared to the catalyst (**IL**), indicating the interaction between iron oxides and the support is weaker in the first sample.

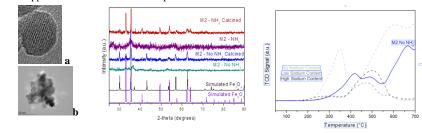


Figure 1 - TEM of catalysts synthesized by PM-Na (top) and IL (bottom), PXRD analysis of ionic liquid method (middle), TPR of ionic liquid method against precipitation method (right).

Significance

We report here on a novel methodology for the controlled synthesis of a Fe₃O₄/HZSM-5 multifunctional catalyst for the direct hydrogenation of CO₂ to gasoline. The product composition and selectivity can be tuned by the choice of by the choice of ionic liquid in the synthetic method and the Fe precursors. This study provides a new pathway for the synthesis of nanocatalysts to produce liquid fuels by utilising CO₂ and H₂, which may in the future lead to alternative approaches to overcome issues with the intermittency of storing and/or utilising energy from renewable sources (photovoltaics, wind energy).

References

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