

# Nanocatalysts from Ionic Liquid Precursors for the Direct Conversion of CO<sub>2</sub> to Hydrocarbons

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## Introduction

The direct conversion of carbon dioxide (CO<sub>2</sub>) into lower olefins (C<sub>2</sub>-C<sub>4</sub>) is a highly desirable process as a sustainable production route. Lower olefins, i.e., ethylene, propylene and butenes (C<sub>2</sub>-C<sub>4</sub>), are key building blocks in the current chemical industry. The reaction proceeds via two main consecutive reactions: Reverse Water Gas Shift (RWGS) to produce CO followed by the further conversion of CO to hydrocarbons via the Fischer-Tropsch reaction<sup>2</sup>. This process is achieved by a multifunctional iron-based catalyst supported on zeolites providing three types of active sites (Fe<sub>3</sub>O<sub>4</sub>, Fe<sub>3</sub>C<sub>2</sub> and acid sites), which cooperatively catalyse a tandem reaction<sup>1</sup>. To date, attempts at synthesising a suitable catalyst for the direct hydrogenation reaction follow a conventional precipitation procedure, whereby Iron Oxide Nanoparticles (IONs) are produced and then embedded within a zeolite structure by granule mixing. This method provides no control over the size and shape of the IONs formed; a characteristic of imperative importance due to its significant effect on the hydrocarbon product distribution obtained. In our novel approach, ionic liquids are utilised for the synthesis of the IONs resulting in better control over size and morphology of the nanostructured material, and as a consequence, better conversion and selectivity towards the olefins.

## Materials and Methods

Na-Fe<sub>3</sub>O<sub>4</sub> nanocatalysts obtained by a one-pot synthesis method employing a precipitation method involving a mixture of iron (II) and iron (III) chloride hydrate, deionised water and HCl to result in a clear solution. Subsequently, NaOH has been added as a precipitating agent, which resulted in the formation of a black precipitate, which consisted of magnetite, Fe<sub>3</sub>O<sub>4</sub>. The compound was then characterised using IR, PXRD (see Fig. 1), SEM / EDX and TPR (Fig. 2). Depending on the amount of washing water, the catalyst contained varying amounts of residual sodium content of up to 7·10<sup>-2</sup> %.

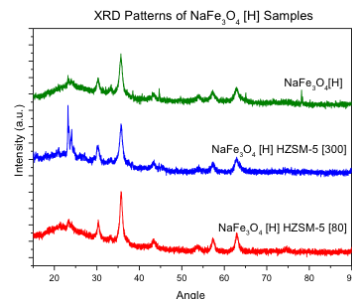
The ionic liquid-assisted synthesis involved heating the reaction medium consisting of the ionic liquid 1-butyl-3-methyl imidazolium bistriflimide, [C<sub>4</sub>mim][Tf<sub>2</sub>N], oleic acid and iron pentacarbonyl under reflux. The precursor iron pentacarbonyl decomposed in a controlled manner by heating the sample up; CO is produced and the iron reacts with residual H<sub>2</sub>O in the ionic liquid mixture to result in Fe<sub>3</sub>O<sub>4</sub>. Following decomposition, the produced magnetite nanoparticles are separated from the reaction medium through application of a neodymium magnet. The Na-Fe<sub>3</sub>O<sub>4</sub>/Zeolite catalyst was typically prepared by granule mixing Na-Fe<sub>3</sub>O<sub>4</sub> catalysts prepared with the methods above with zeolites (HZSM(SiO<sub>2</sub>/Al<sub>2</sub>O<sub>3</sub> = 80 and 300) in a ball mill at a mass ratio of the two components of 1:1.

CO<sub>2</sub> hydrogenation reactions were performed at 320 °C, 3 MPa H<sub>2</sub>/CO 3,3 in a stainless steel fixed-bed reactor with an inner diameter of 15 mm. Typically, 1 g of composite catalyst (20–40 meshes) with Na-Fe<sub>3</sub>O<sub>4</sub>/Zeolite 1/4 1/1 (mass ratio) was used. Prior to reaction, the catalyst was in-situ reduced at 350 °C for 8 h in a pure H<sub>2</sub> flow at atmospheric pressure.

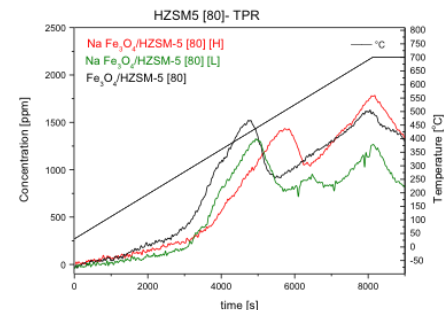
All of the products from the reactor were introduced in a gaseous state and analysed with an online gas chromatograph (GC).

## Results and Discussion

The ionic-liquid assisted synthesis of a nanocrystalline magnetite precursor showed that ionic liquids provide a controlled precipitation method thanks to their dual functionality as solvent and templating agent. This level of control over the morphology of the produced IONs allows for the selectivity of the hydrocarbon distribution to be directly tailored to light olefin production. Characterization of the prepared catalysts by PXRD (Figure 1) shows phase purity for the magnetite obtained from the conventional precipitation method, and shows no alteration after granulation with the zeolite, showing the presence of high purity in Fe<sub>3</sub>O<sub>4</sub>, small particle size and good dispersion with the zeolite component. The compounds obtained by the ionic liquid methods result in much broader XRD patterns, pointing at less crystallinity and smaller particle size. This has also been confirmed by SEM and TEM. Hydrogen temperature-programmed reduction (H<sub>2</sub>-TPR) was used to determine the interaction between Fe species and the support. As shown in Figure 2 for the HZSM 80, all the catalysts present two peaks with increasing reduction temperature, which are assigned to the conversions Fe<sub>3</sub>O<sub>4</sub>→FeO and FeO→Fe, respectively. It is observed that Fe<sub>3</sub>O<sub>4</sub>/HZSM80 start to be reduced at lower temperature compared to the catalyst with low (L) and high (H) Na content. This indicates the interaction between iron oxides and the support is weaker if Na is completely removed.



**Figure 1.** XRD patterns of Na-Fe<sub>3</sub>O<sub>4</sub> sample with high Na content (H) and Na-Fe<sub>3</sub>O<sub>4</sub> mixed with HZSM with SiO<sub>2</sub>/Al<sub>2</sub>O<sub>3</sub> = 80 and 300.



**Figure 2.** TPR profile of 4000 ppm H<sub>2</sub> in Ar from r.t. to 700°C, 5°C/min

## Significance

We report here on a novel methodology for the controlled synthesis of a Na-Fe<sub>3</sub>O<sub>4</sub>/HZSM-5 multifunctional catalyst for the direct hydrogenation of CO<sub>2</sub> to gasoline. The catalytic testing under industrially relevant conditions resulted in improved selectivity to C<sub>5</sub>-C<sub>11</sub> as well as low CH<sub>4</sub> and CO selectivity. Furthermore, the product composition can be tuned by the choice of zeolite type, and by the choice of ionic liquid in the synthetic method. This study provides a new pathway for the synthesis of nanocatalysts for the production of liquid fuels by utilising CO<sub>2</sub> and H<sub>2</sub>, which may in the future lead to alternative approaches to overcome issues with the intermittency of storing and/or utilising energy from renewable sources (photovoltaics, wind energy).

## References

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